Biological and Mechanical Evaluation of Novel Prototype Dental Composites

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Appendix

Calculating Oligomer formulation

The thiol–ene polymerization follows a step-growth mechanism, and as a result the gelation point of this system can be predicted. Following the Flory-Stockmayer theory, the gelation point of multifunctional monomer systems can be predicted according to the following equation.

$$\partial = \frac{1}{\sqrt{r(f_a - 1)(f_b - 1)}}$$

In which f_a and f_b are the functionalities of the 2 monomers, r is the molar ratio of the monomers ($r = N_a/N_b \pm 1$), and α is the extend of reaction at which gelation occurs. α was set to be ~1.02, so that the maximum amount of oligomerization would occur, while maintaining a flowable oligomer. For the TMES/TNTATO system, $f_a = 3$, and $f_b = 4$, r was calculated to be 0.16. For TMES/DNBPA, $f_a = 2$, and $f_b = 4$, resulting in r = 0.3.



Appendix Figure 1. Structures of monomers used in this study. a) 1,3,5-Triallyl-1,3,5-triazine-2,4,6-trione (TATATO), b) Tetra(2-mercaproethyl)silane (TMES), c) trinorbornyl triazine trione (TNTATO), d) dinorbornyl bisphenol A (DNBPA), e) Triethyleneglycol dimethacrylate (TEGDMA), f) bisphenol A-glycidyl methacrylate (bisGMA)



Appendix Figure 2. Synthesis of TMES from tetravinylsilane and thioacetic acid followed by hydrolosis.



Appendix Figure 3. Synthesis of TNTATO from TATATO and dicyclopentadiene.

 l_{\sim} 190°C $\widehat{\left(\right)}$ <u>`0′</u> \cap

Appendix Figure 4. Synthesis of DNBPA from Diallyl bisphenol A and dicyclopentadiene.



Appendix Figure 5. Oligomerization reaction of TMES with TNTATO in Tetrahydrofuran.



Appendix Figure 6. Oligomerization reaction of TMES with DNBPA in Tetrahydrofuran.



Appendix Figure 7. Time/conversion plots of unfilled TMES-TNTATO/TATATO (black, squares), and TMES-DNBPA/TATATO (red, circles) formulated with 1% TPO and irradiated at 405 nm with 10 mW/cm². Conversions are obtained from monitoring the TATATO C=C in the FTIR spectrum at 3083 cm⁻¹.